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easily ~~Equilibrium~~

~~2~~ ~~Calculating~~

~~Equilibrium~~ *Which way*

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Shift? (Le Chatelier's

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View Answer. The

equilibrium system

below shows the

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reaction of sour gas with methane to produce carbon disulfide and hydrogen gas. $\text{CH}_4(\text{g}) + 2\text{H}_2\text{S}(\text{g}) \rightarrow \text{CS}_2(\text{g}) + 4\text{H}_2(\text{g})$; $\Delta H = +232.5 \text{ kJ}$...

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10^{-3} and 10^3 .

Chemical equilibrium (definition) The state

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reached when the concentration of reactants and products remain constant over time. What is the general, reversible reaction for equilibrium? (With letters) $aA + bB \rightleftharpoons cC + dD$. Equilibrium equation (general with letters) $K_c = \frac{[C]^c[D]^d}{[A]^a[B]^b}$.

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Equilibrium means: 1.

Stable state from where
neither the seller nor the
buyer wants to deviate.

2. It is the intersection
point of the market
supply and demand

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curves. Equilibrium
means the quantity
supplied equals ... -

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26 D5 Chemical

equilibrium is said to be dynamic because A. the reaction proceeds quickly. B. the mass of the reactants is decreasing. C. the macroscopic properties are constant. D. both forward and reverse reactions are occurring.

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27 D5 Equilibrium is a dynamic process because the A. macroscopic properties are not changing.

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system at equilibrium ...
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carefully the Sample
Problem, “Using Le
Châtelier’s Principle”
on page 528 before you

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do questions 2.5 and 2.6

|| **See your instructor
to find out what

questions you should

complete for review. 2.3

(a) State in which

direction the

equilibrium will shift for

each of the following,

when there are

~~Equilibrium/Acids and~~

~~Bases Study Guide~~

This model should show

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the three states that occur in the chemical equilibrium process. As students work, walk around to offer suggestions and guide. Have students:

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the stress. 2. Answers
may vary, but should

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include changes in
concentration, volume,
and temperature. Some
students may list
pressure in place of
volume. 3. left; remove
N₂ or H₂

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Constant: A reversible
chemical reaction
system can proceed in
both directions from a
set of initial conditions
(concentrations) towards
a state of chemical
equilibrium.

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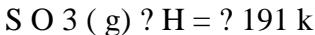
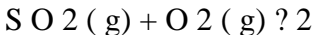
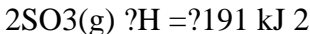
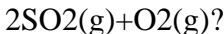
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~~Answers~~
The given reaction

equation is:



J. The equilibrium

constant is a ratio of the

forward and reverse ...

~~Consider the following~~

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~~equilibrium reaction. 2~~

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1) Give a balanced chemical equation which describes this equilibrium. 2) Explain what influence the removal of H_2 from the container will have on the equilibrium. 3) Explain what effect...

~~You are given a container in which H_2 (~~

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Solved: Consider the following equilibrium:

$4\text{HCl}(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{H}_2\text{O}(\text{g}) + 2\text{Cl}_2(\text{g}); K_c = 357$. An equilibrium mixture of the system was found to...

~~Consider the following equilibrium: $4\text{HCl}(\text{g})$~~
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Changing the temperature of a

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reaction at equilibrium
alters both the
equilibrium. (12)and the
equilibrium position.

When a reaction is. (13),
which means it releases
energy, lowering the
temperature shifts the
equilibrium to the
(14)because the forward
reaction liberates heat
and removes the (15).

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knowledge but also to help students get inspired to explore and discover many creative ideas from themselves.

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and numerical problem

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