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---

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#Chapter 4 #OTQs

#MCQs #1 Mark

Questions Chemistry



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Correct Ans. in 10 to  
20 sec. Some

~~Important Points For~~

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Thermodynamics -

very important

questions sheet

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kinetics | asked in ssc

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~~PHARMACY |~~

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(Multiple Choice

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Rate Laws –

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/u0026 Equations

Chemical kinetics

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MCQs | Arvind Sir

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Questions. MCQs on  
Chemical Kinetics : 1.~~

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Kinetics  
Multiple Choice  
Questions And  
Answers

The rate of a chemical reaction tells us about (A) the reactants taking part in the reaction (B) the products formed in the reaction (C) how slow or fast the reaction is taking place

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Chemical Kinetics  
and Answers~~

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Chemical

Kinetics

Multiple Choice

Questions Answers:

Chemical Kinetics

Examples Questions

Question 1. The rate constant is doubled when temperature increases from  $27^{\circ}\text{C}$  to  $37^{\circ}\text{C}$ . Activation energy in kJ is. A. 54. B. 100. C. 34. D. 50.

Question 2. A catalyst increases the rate of



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reaction because it ...

Multiple Choice

~~Chemical Kinetics~~

~~Exam Questions with~~

~~Answers - NEET ...~~

Chemical Kinetics:

Multiple choice

questions with

answers. Chemistry :

Chemical Kinetics:

Multiple choice

questions with

answers, Solution

and Explanation.

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EVALUATION . 1. For a first order reaction A → B the rate constant is  $k \text{ min}^{-1}$  ...

Answers

~~Chemical Kinetics:  
Multiple choice  
questions with  
answers~~

The following section consists of Chemistry Multiple Choice questions on Chemical Kinetics For

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Competitions and

exams. Select the  
correct option to test  
your skills on

Chemical Kinetics. Set

1

~~MCQ on Chemical~~

~~Kinetics - Quiz and~~

~~MCQ Preparations for~~

~~...~~

A.P. Chemistry

Practice Test: Ch. 12,

Kinetics MULTIPLE

*Page 19/116*

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CHOICE. Choose the one alternative that best completes the statement or answers the question. 1)

Consider the following reaction:  
 $3A \rightarrow 2B$  The average rate of appearance of B is given by  $D[B]/Dt$ . Comparing the rate of appearance of B and the rate of

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~~A.P. Chemistry~~

~~Practice Test: Ch. 12,  
Kinetics MULTIPLE ...~~

~~Questions And~~  
Rates of Reaction:

~~Answers~~  
Chemical Kinetics 50

I. Multiple Choice 20

1. The rate

determining step for  
a complex reaction is

the one which is A.

fastest C. slowest B.

last in the sequence

D. first in the

sequence 2. Which

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Chemical

one of the following  
is NOT a key concept  
of the collision  
theory: A. particles  
must collide in order  
to react

~~I. Multiple Choice 20~~  
Chemical Kinetics  
MCQs with answers,  
Test: 1, Total  
Questions: 15

~~Chemical Kinetics~~

*Page 22/116*

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~~MCQ | Questions~~

~~Paper 1~~

Kinetics: Multiple

Choice Review

Questions 1. In the

rate law, Rate =

$k[\text{NO}]^2[\text{O}_2]$  ... The

slow rate of a

particular chemical

reaction might be

attributed to which

of the following? (a) a

low activation energy

... Use the kinetics

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kinetics provided in the  
previous problem to  
answer this question.

What is the rate of

Answers

~~Kinetics: Multiple  
Choice Review  
Questions~~

Test prep MCAT  
Chemical processes  
Kinetics. Kinetics.

Practice: Kinetics  
questions. This is the  
currently selected



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Chemical

item. Rate of

reaction. Rate law  
and reaction order.

Experimental

determination of rate

laws. First-order

reaction (with

calculus) Plotting

data for a first-order

reaction.

~~Kinetics questions~~

~~(practice) | Kinetics |~~

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(a) the same as its balanced chemical equation. (b) obvious if its heat of reaction is known. (c) obvious if its reaction order is known. (d) sometimes difficult to prove. (e) obvious if its activation energy is known. 17.

Suppose the activation energy of a certain reaction is

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250 kJ/mol.

Multiple Choice

~~Sample Questions~~

~~Chapter 16~~

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~~Multiple Choice~~

Questions (MCQ) for  
Chemical Kinetics -  
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extremely critical for  
all CBSE students to  
score better marks.

~~Multiple Choice~~

~~Questions (MCQ)~~

~~Chemical Kinetics ...~~

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Multiple Choice

Questions (Type-II)

Note : In the

following questions

two or more options

may be correct. Rate

law cannot be

determined from

balanced chemical

equation if \_\_\_\_\_. (i)

reverse reaction is

involved. (ii) it is an

elementary reaction.

(iii) it is a sequence of

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elementary reactions.

(iv) any of the reactants is in excess.

~~Class 12 Important~~

~~Questions for~~

~~Chemistry —~~

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~~MCQ Questions for~~

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~~Chapter 4 Chemical...~~

Multiple Choice  
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Kinetics Class 12

Important Questions

Very Short Answer

Type Question 1.

Define ' rate of a  
reaction ' . (Delhi

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2010) Answer: Rate of a reaction: Either, The change in the concentration of any one of the reactants or products per unit time [...]

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...

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and Answers of

Chapter 12: Chemical  
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Questions And

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Kinetics Practice

Problems For

multiple-choice

questions, show your

work below the a-e

choices for full credit

AND circle your

answer. 1. Calculate

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the average speed of  
reaction if at 10.0  
seconds the reactant  
concentration is  
1.25M but at 55.0  
seconds the  
concentration is  
1.02M. a.

~~(Hint: first find k)~~

Chapter 12: Multiple  
choice questions.  
Instructions. Answer  
the following

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Questions and then  
press 'Submit' to get  
your score. Question  
1 A drug suspension  
decomposes by zero-  
order kinetics with a  
rate constant of  $2 \text{ mg mL}^{-1} \text{ month}^{-1}$ . If the  
initial concentration  
is  $100 \text{ mg mL}^{-1}$ , what  
is the shelf life ( $t_{10\%}$ )?  
a) 2 months

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Multiple Choice  
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chemistry of life,  
electrode potential,  
electrons in atoms,  
enthalpy change,  
equilibrium, group IV,  
groups II and VII,  
halogenoalkanes,  
hydrocarbons,  
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organic chemistry,

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lattice energy, moles  
and equations,  
nitrogen and sulfur,  
organic and nitrogen  
compounds,  
periodicity,  
polymerization, rates  
of reaction, reaction  
kinetics, redox  
reactions and  
electrolysis, states of  
matter, transition  
elements tests for

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Compound MCQs

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Compounds MCQs

Chapter 5: Carboxylic  
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and Sulfur MCQs  
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Chapter 27: States of

Matter MCQs Chapter

28: Transition

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alcohols, and  
alcohols reactions.

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questions: Atom

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facts, elements and

atoms, number of

nucleons, protons,

electrons, and

neutrons. Solve

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questions:

Introduction to

benzene, arenes

reaction, phenol and

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properties, and  
reactions of phenol.  
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carbonyl compounds,  
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with HCN,

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kinetics

iodomethane. Solve  
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questions: Chemical  
bonding types,  
chemical bonding  
electron pair, bond  
angle, bond energy,  
bond energy, bond  
length, bonding and  
physical properties,

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Kinetics energy,  
repulsion theory,  
Multiple Choice  
Questions And  
Answers  
double covalent  
bonds, triple covalent  
bonds, electron pair  
repulsion and bond  
angles, electron pair  
repulsion theory,  
enthalpy change of  
vaporization,  
intermolecular forces,  
ionic bonding, ionic

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bonds and covalent bonds, ionic bonds, metallic bonding, metallic bonding and delocalized electrons, number of electrons, sigma bonds and pi bonds, sigma-bonds, pi-bonds, s-orbital and p-orbital, Van der Waals forces, and contact points. Solve "Chemistry of Life MCQ" PDF book with



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specifity, enzymes,  
reintroducing amino  
acids, and proteins.

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electrolysis process,

measuring standard

electrode potential,

quantitative

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and oxidation. Solve

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evidence, ionization

energy, periodic

table, simple

electronic structure,

sub shells, and

atomic orbitals. Solve

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bond energies,

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introduction to

energy changes,

measuring enthalpy

changes. Solve

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Equilibrium constant

expression,

equilibrium position,

acid base equilibria,

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chemical industry  
equilibria, ethanoic  
acid, gas reactions  
equilibria, and  
reversible reactions.

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IV, metallic character  
of group IV elements,  
ceramic, silicon oxide,

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Answers

covalent bonds,  
properties variation  
in group IV, relative  
stability of oxidation  
states, and tetra  
chlorides. Solve  
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number of group II  
metals, covalent  
bonds, density of

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physical properties of  
group II elements,  
physical properties of  
group VII elements,  
reaction of group II  
elements with  
oxygen, reactions of  
group II elements,  
reactions of group VII  
elements, thermal  
decomposition of  
carbonates and  
nitrates, thermal  
decomposition of



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group II carbonates,  
thermal  
decomposition of  
group II nitrates, uses  
of group II elements,  
uses of group II  
metals, uses of  
halogens and their  
compounds. Solve  
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halogenoalkanes,  
elimination reactions,  
nucleophilic  
substitution in  
halogenoalkanes,  
and nucleophilic  
substitution  
reactions. Solve  
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alkanes, sources of  
alkanes, addition  
reactions of alkenes,  
alkane reaction,  
alkenes and formulas.

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chemistry, functional

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solutions, equilibrium

and solubility,

indicators and acid

base titrations, pH

calculations, and

weak acids. Solve

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Introduction to

lattice energy, ion

polarization, lattice

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energy value,  
atomization and  
electron affinity, Born  
Haber cycle, and  
enthalpy changes in  
solution. Solve

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substance, atoms,  
molecules mass,  
chemical formula and

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volumes, mole

calculations, relative

atomic mass,

solutions, and

concentrations. Solve

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compounds, nitrogen

and gas properties,

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ammonium

compounds,

environmental

problems caused by

nitrogen compounds

and nitrate fertilizers,

sulfur and oxides,

sulfuric acid and

properties, and uses

of sulfuric acid. Solve

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Nitrogen Compounds

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amino acids, peptides  
and proteins. Solve

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questions: Acidic

oxides, basic oxides,

aluminum oxide,

balancing equation,

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period 3 chlorides,  
balancing equations:  
reactions with  
chlorine, balancing  
equations: reactions  
with oxygen,  
bonding nature of  
period 3 oxides,  
chemical properties  
of chlorine, chemical  
properties of oxygen,  
chemical properties  
periodicity, chemistry  
periodic table,

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Chemistry: oxides,  
chlorides of period 3  
elements, electrical  
conductivity in  
period 3 oxides,  
electronegativity of  
period 3 oxides, ionic  
bonds, molecular  
structures of period 3  
oxides, oxidation  
number of oxides,  
oxidation numbers,  
oxides and  
hydroxides of period

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3 elements, oxides of  
period 3 elements,  
period III chlorides,  
periodic table  
electronegativity,  
physical properties  
periodicity, reaction  
of sodium and  
magnesium with  
water, and relative  
melting point of  
period 3 oxides. Solve  
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polymerization,

polyamides,

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theory, effect of

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of reaction. Solve  
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physical properties of  
transition elements,  
redox and oxidation.

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who are weak in

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chemistry or have no chemistry on entry to university. Such students usually find physical chemistry the most difficult part of the chemistry course, and within this section numerical problem solving is an additional difficulty. The text should also be invaluable to first

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kinetics  
chemists. This text  
provides an  
introduction to  
physical chemistry  
and the gas laws,  
followed by chapters  
on thermodynamics,  
chemical equilibrium,  
electrochemistry and  
chemical kinetics.  
Each section involves  
a brief introduction  
followed by a

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representative

examination

question, which is  
broken down into a

proposed working

method. Both short

multiple-choice

questions and related

full examination-type

questions are

included. This book

will prove invaluable

to students who need

encouragement in a

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logical approach to problem solving in physical chemistry, teaching them to think for themselves when faced with a problem.

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electrode potential,  
electrons in atoms,  
enthalpy change,  
equilibrium, group IV,  
groups II and VII,  
halogenoalkanes,  
hydrocarbons,  
introduction to  
organic chemistry,  
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lattice energy, moles  
and equations,  
nitrogen and sulfur,  
organic and nitrogen

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Compounds,  
periodicity,  
polymerization, rates  
of reaction, reaction  
kinetics, redox  
reactions and  
electrolysis, states of  
matter, transition  
elements to enhance  
teaching and  
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the syllabus of many  
competitive papers  
for admission exams  
of different

universities from  
chemistry textbooks  
on chapters: Alcohols  
and Esters Multiple  
Choice Questions: 27  
MCQs Atomic  
Structure and Theory  
Multiple Choice  
Questions: 37 MCQs  
Benzene: Chemical

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Compound Multiple  
Choice Questions: 41  
MCQs Carbonyl  
Compounds Multiple  
Choice Questions: 29  
MCQs Carboxylic  
Acids and Acyl  
Compounds Multiple  
Choice Questions: 27  
MCQs Chemical  
Bonding Multiple  
Choice Questions:  
213 MCQs Chemistry  
of Life Multiple

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Choice Questions: 29

MCQs Electrode

Potential Multiple

Choice Questions: 62

MCQs Electrons in

Atoms Multiple

Choice Questions: 53

MCQs Enthalpy

Change Multiple

Choice Questions: 45

MCQs Equilibrium

Multiple Choice

Questions: 50 MCQs

Group IV Multiple

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Chemical

Choice Questions: 53

MCQs Groups II and

VII Multiple Choice

Questions: 180 MCQs

Halogenoalkanes

Multiple Choice

Questions: 33 MCQs

Hydrocarbons

Multiple Choice

Questions: 53 MCQs

Introduction to

Organic Chemistry

Multiple Choice

Questions: 52 MCQs

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Chemical

Ionic Equilibria

Multiple Choice

Questions: 56 MCQs

Lattice Energy

Multiple Choice

Questions: 33 MCQs

Moles and Equations

Multiple Choice

Questions: 50 MCQs

Nitrogen and Sulfur

Multiple Choice

Questions: 89 MCQs

Organic and Nitrogen

Compounds Multiple

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Chemical

Choice Questions: 54

MCQs Periodicity

Multiple Choice

Questions: 202 MCQs

Polymerization

Multiple Choice

Questions: 36 MCQs

Rates of Reaction

Multiple Choice

Questions: 39 MCQs

Reaction Kinetics

Multiple Choice

Questions: 52 MCQs

Redox Reactions and



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Electrolysis Multiple  
Choice Questions: 55  
MCQs States of  
Matter Multiple  
Choice Questions: 66  
MCQs Transition  
Elements Multiple  
Choice Questions: 29  
MCQs The chapter  
"Alcohols and Esters  
MCQs" covers topics  
of introduction to  
alcohols, and  
alcohols reactions.

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The chapter "Atomic Structure and Theory MCQs" covers topics of atom facts, elements and atoms, number of nucleons, protons, electrons, and neutrons. The chapter "Benzene: Chemical Compound MCQs" covers topics of benzene, arenes reaction, phenol properties, and

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reactions of phenol.

The chapter

"Carbonyl

Compounds MCQs"

covers topics of

carbonyl compounds,

aldehydes and

ketone testing,

nucleophilic addition

with HCN,

preparation of

aldehydes and

ketone, reduction of

aldehydes, and

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